

Balancing Equations with Polyatomic Ions

- $\text{FeCl}_3 + (\text{NH}_4)_2\text{S} \longrightarrow \text{Fe}_2\text{S}_3 + \text{NH}_4\text{Cl}$
- $\text{Al}(\text{OH})_3 + \text{HNO}_3 \longrightarrow \text{Al}(\text{NO}_3)_3 + \text{H}_2\text{O}$
- $\text{Ba}(\text{OH})_2 + \text{Fe}_2(\text{SO}_4)_3 \longrightarrow \text{Fe}(\text{OH})_3 + \text{BaSO}_4$
- $\text{NH}_4\text{Cl} + \text{Na}_3\text{PO}_4 \longrightarrow (\text{NH}_4)_3\text{PO}_4 + \text{NaCl}$
- $\text{Mg}(\text{OH})_2 + \text{H}_3\text{PO}_4 \longrightarrow \text{Mg}_3(\text{PO}_4)_2 + \text{H}_2\text{O}$
- $\text{Na}_2\text{CO}_3 + \text{H}_3\text{PO}_4 \longrightarrow \text{Na}_3\text{PO}_4 + \text{H}_2\text{O} + \text{CO}_2$
- $(\text{CH}_3\text{COO})_2\text{Pb} + (\text{NH}_4)_2\text{S} \longrightarrow \text{PbS} + \text{CH}_3\text{COONH}_4$
- $\text{NaI} + \text{MnO}_2 + \text{H}_2\text{SO}_4 \longrightarrow \text{Na}_2\text{SO}_4 + \text{MnSO}_4 + \text{H}_2\text{O} + \text{I}_2$
- $\text{KMnO}_4 + \text{H}_2\text{SO}_3 \longrightarrow \text{K}_2\text{SO}_4 + \text{MnSO}_4 + \text{H}_2\text{SO}_4 + \text{H}_2\text{O}$
- $\text{Al} + \text{H}_2\text{SO}_4 \longrightarrow \text{Al}_2(\text{SO}_4)_3 + \text{H}_2\text{O} + \text{SO}_2$
11. Lead (II) perchlorate + ammonium iodide \longrightarrow
12. Lithium carbonate + copper (I) nitrate \longrightarrow
13. Ammonium dichromate + iron (II) nitrate \longrightarrow
14. Iron (III) acetate + ammonium sulphide \longrightarrow
15. Ammonium phosphate + aluminium chloride \longrightarrow
16. Sodium sulphate + silver nitrate \longrightarrow
17. Sodium hydrogen carbonate + calcium perchlorate \longrightarrow
18. Copper (II) chloride + potassium phosphate \longrightarrow
19. Barium hydroxide_(aq) + phosphoric acid_(aq) \longrightarrow barium phosphate_(aq) + water_(l)
20. Calcium nitrate_(aq) + ammonium oxalate_(aq) \longrightarrow calcium oxalate_(s) + ammonium nitrate_(aq)