Solutions:

Homogenous: constant compositi0n
Solute: substance being dissolved (smaller quantity, usually)
Solvent: substance doing dissolving
Solubility: maximum amount of solute that can be dissolved at a given
temp. in 100g of the solvent.
~ SOLIDS → increases as temp. increases (generally)
~ GASES → decreases as temp. increases.

Formulas:
1. n= cv   # of moles (mol) = concentration(mol/L) x volume (L)
- IB uses dm^3 instead of L $(1L = 1dm^3)$
- Higher the concentration, higher the molarity

2. mass % = $\underline{\text{mass of solute (g)}}$ x 100
          mass of solute + mass of solvent

3. volume % = $\underline{\text{volume of solute}}$ x 100
          solume of solution

4. day,

making solutions:
ex. Calculate the volume of water required and calculate the solution required when
5.5% beer is diluted to a high glass container with 100mL of water, diluted into 1.25%

stock	diluted:
c.= 5.5%	c = 1.25%
v = ?	v = 100mL

$C_1V_1 = C_2V_2$ the volume of water in the flask is:
$V_1 = \underline{(1.25)(100)}$ 100 - 22.7 = 77.3mL
          5.5%
$\boxed{= 22.7\text{mL of stock beer}}$

mass of gravimetric analysis

problems: