

## Review of Polyatomic Compounds

Write correct formulas of the compounds formed when the positive ions, (i.e. cations), in the vertical column combine with the negative ions, (i.e. anions), listed across the top row.

A few are done for you.

Cation	nitrate $\text{NO}_3^{-1}$	hydroxide $\text{OH}^{-1}$	bicarbonate $\text{HCO}_3^{-1}$	chlorate $\text{ClO}_3^{-1}$	carbonate $\text{CO}_3^{-2}$	sulphate $\text{SO}_4^{-2}$	phosphate $\text{PO}_4^{-3}$
Lithium, $\text{Li}^{+1}$	$\text{LiNO}_3$						
Calcium		$\text{Ca(OH)}_2$					
Aluminium					$\text{Al}_2(\text{CO}_3)_3$		
Tin (II)						$\text{SnSO}_4$	
Arsenic(III)			$\text{As(HCO}_3)_3$				
Ferrous	$\text{Fe(NO}_3)_2$						
Chromium(III)							
Cupric				$\text{Cu(ClO}_3)_2$			
silver							$\text{Ag}_3\text{PO}_4$
Zinc		$\text{Zn(OH)}_2$					
Nickel (II)							$\text{Ni}_3(\text{PO}_4)_2$
Magnesium							
Plumbous	$\text{Pb(NO}_3)_2$						
Mercury (II)							
Stannic					$\text{Sn(CO}_3)_2$		